



## Chemistry

Time Remaining: 45/45 (Minutes)

Q.1

Test 5 Solids

CHEMISTRY NMDCAT

A crystal 'x' has very high melting point and is totally insoluble in water and does not conduct electricity is likely to be \_\_\_\_\_ crystal:

- a. Ionic
- b. Covalent
- c. Molecular
- d. Metallic

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Correct Answer:

☐ A ☐ B ☐ C ☐ D

Next



Time Remaining: 44/45 (Minutes)

Q.2

Test 5 Solids

CHEMISTRY NMDCAT

Dry Ice is solid:

- a.  $\text{SO}_2$
- b.  $\text{CO}_2$
- c. CO
- d.  $\text{O}_2$

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.3

Test 5 Solids

CHEMISTRY NMDCAT

The inter ionic distance is maximum in the crystal lattice of

a. LiCl

b. NaCl

c. RbCl

d. KCl

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Correct Answer:



A



B



C



D

Next

Back





Time Remaining: 44/45 (Minutes)

Q.4

Test 5 Solids

CHEMISTRY NMDCAT

Among the properties which property correctly match with metallic solids:

- i. They show metallic lusters
  - ii. Their conductance decreases with the rise of temperature
  - iii. Malleable and ductile
  - iv. Cubic and hexagonal close packing
- a. i,ii and iv                      b. i,ii and iii  
c. i ,ii, iii and iv                d. I,iii and iv

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Correct Answer:

☒ A   ☐ B   ☐ C   ☐ D

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Back



Time Remaining: 44/45 (Minutes)

Q.5

Test 5 Solids

CHEMISTRY NMDCAT

**Which statement about covalent solids is incorrect?**

- a. They contain a network of atoms
- b. They have high melting points
- c. They are very hard and greater energy is required to break them
- d. They volatility is very high

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Correct Answer:



A



B



C



D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.6

Test 5 Solids

CHEMISTRY NMDCAT

**Covalent crystals are bad conductor of electricity due to absence of free electrons and ions except:**

- a. Silicon Carbide
- b. Graphite
- c. Cadmium iodide
- d. Born nitride

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back





Time Remaining: 44/45 (Minutes)

Q.7

Test 5 Solids

CHEMISTRY NMDCAT

**The covalent crystals having giant molecules like diamond and silicon carbide are:**

- a. Soluble in all the solvents
- b. Insoluble in all the polar solvents only
- c. Soluble in all the non-polar solvents only
- d. Insoluble in all the solvents

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.8

Test 5 Solids

CHEMISTRY NMDCAT

In diamond each carbon atom is:

- a. sp-hybridized
- b.  $sp^2$ -hybridized
- c.  $sp^3$ -hybridized
- d. unhybridized

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

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Time Remaining: 44/45 (Minutes)

Q.9

Test 5 Solids

CHEMISTRY NMDCAT

The overall structure of diamond looks like:

- a. Body centered cubic
- b. Face centered cubic
- c. End centered
- d. none of given

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Correct Answer:



A



B



C



D

Next

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Time Remaining: 44/45 (Minutes)

Q.10

Test 5 Solids

CHEMISTRY NMDCAT

Which one of the following pairs contains polar molecular solids?

- a. Iodine and Sugar
- b. Carbon dioxide and Ice
- c. Phosphorus and Carbon dioxide
- d. Sugar and Ice

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

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Time Remaining 44/45 (Minutes)

Q.11

Test 5 Solids

CHEMISTRY NMDCAT

Valence bond theory treat metallic bond as:

- a. A coordinate covalent bond
- b. A localized covalent bond
- c. A de-localized covalent bond
- d. Van der Waal's force

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Don't Know



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Time Remaining 44/45 (Minutes)

QUIZ

Test 5 Solids

CHEMISTRY NMDCAT

**Silicon carbide forms**

- a. ionic solid crystal
- b. covalent crystal
- c. molecular crystal
- d. b, c

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Correct Answer:

1 2 3 4 5 6 7 8 9 10 11 12

Next

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Time Remaining 43/45 (Minutes)

QUIZ

Test 5 Solids

CHEMISTRY NMDCAT

Substance having m.p higher than  $500^{\circ}\text{C}$  and insoluble in  $\text{H}_2\text{O}$  and organic solvent and conductor in solid and phase:

- |           |                    |
|-----------|--------------------|
| a. Copper | b. Sodium chloride |
| c. Silica | d. Cell            |

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0/12 100% Made Up 0/100 0/100

Correct Answer:

0/4 0/11 0/12 0/13

Next

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Time Remaining 43/45 (Minutes)

Q111

Test 5 Solids

CHEMISTRY NMDCAT

The crystalline solid which show very slow chemical reactions

- a. ionic  
c. molecular  
c. covalent  
d. b, c

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Correct Answer:

4 11 12 13

Next

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Time Remaining 43:45 (Minutes)

QUIZ

Test 5 Solids

CHEMISTRY NMDCAT

The arrangement ABC, ABC .... is referred as

- a. cubic close packing
- b. octahedral closed packing
- c. hexagonal closed packing
- d. tetrahedral closed packing

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012 345 678 9012 345 678 9012

Don't Worry

1 2 3 4 5 6 7 8 9 10

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Back



Time Remaining 43:45 (Minutes)

Q.18

Test 5 Solids

CHEMISTRY NMDCAT

**Which statement is not true about metallic solid?**

- a. They are ductile and malleable
- b. They are conductor of electricity
- c. their conductivity increases by increasing temperature
- d. they are lustrous

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1 2 3 4 5 6 7 8 9 10

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Time Remaining 43:45 (Minutes)

QUIZ

Test 5 Solids

CHEMISTRY NMDCAT

Which solid does not contain covalent bonds

- a. Copper
- b. Diamond
- c. Graphite
- d. Ice

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Don't Know

1 2 3 4 5 6 7 8 9 10

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Time Remaining 43:45 (Minutes)

QUIZ

Test 5 Solids

CHEMISTRY NMDCAT

The branch of science which deals with the study of the structure of crystal is called

- a. Chymography
- b. Crystallography
- c. Chromatography
- d. Industrial chemistry

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Correct Answer

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Time Remaining 43:45 (Minutes)

Q.11

Test 5 Solids

CHEMISTRY NMDCAT

Glass may begin to crystallize by a process called

- a. annealing
- b. etching
- c. distillation
- d. none of these

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Correct Answer:

1 2 3 4 5 6 7 8 9 10 11 12 13 14 15 16 17 18 19 20 21 22 23 24 25 26 27 28 29 30 31 32 33 34 35 36 37 38 39 40 41 42 43 44 45 46 47 48 49 50 51 52 53 54 55 56 57 58 59 60 61 62 63 64 65 66 67 68 69 70 71 72 73 74 75 76 77 78 79 80 81 82 83 84 85 86 87 88 89 90 91 92 93 94 95 96 97 98 99 100

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Time Remaining 43/45 (Minutes)

0.30

Test 5 Solids

CHEMISTRY NMDCAT

The crystalline part of other wise amorphous solids is called

- |                    |                |
|--------------------|----------------|
| a. crystal system  | b. crystallite |
| c. crystal lattice | d. allotrope   |

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Time Remaining 43:45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

Which substance has diffused melting point

- a. crystalline solid
- b. amorphous solid
- c. metallic solid
- d. covalent solid

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Time Remaining 43/45 (Minutes)

Q.12

Test 5 Solids

CHEMISTRY NMDCAT

Which has the strongest bonding in the solid state?

- a. Hydrogen Chloride (HCl)
- b. Chlorine (Cl<sub>2</sub>)
- c. Xenon(Xe)
- d. Sodium Chloride (NaCl)

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Q.12

100%

100%

100%

100%

Correct Answer:



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Time Remaining 42/45 (Minutes)

Q.11

Test 5 Solids

CHEMISTRY NMDCAT

**When substance moves from a solid to a liquid state all of the following changes occur except**

- a. Molecules become more disordered
- b. K.E of the molecules decreases
- c. Intermolecular forces become weaker
- d. Molecule become further separated

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0/12 100% Image Description

Correct Answer:

0/4 100% 100% 100%

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Time Remaining 42/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

**When the atoms of the third layer are arranged in such a way that they directly lie above the atoms of the first layer then this arrangement is called**

- a. ABAB (hexagonal)
- b. ABCABC (Cubic)
- c. Orthorhombic
- d. Rhombohedral

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Don't Know



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Time Remaining 42/45 (Minutes)

Q.38

Test 5 Solids

CHEMISTRY NMDCAT

Which one is a conductor but is not malleable?

- a. Iron
- b. Graphite
- c. Silver
- d. Platinum

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Correct Answer:

0/4 100% 100% 100%

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Time Remaining 42/45 (Minutes)

Q38

Test 5 Solids

CHEMISTRY NMDCAT

**A malleable solid is one which can be**

- a. Converted into wires
- b. Converted into thin sheets
- c. Melted easily
- d. All of the above

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

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Time Remaining 42/45 (Minutes)

Q17

Test 5 Solids

CHEMISTRY NMDCAT

**Buckyballs is an allotropic form of**

- a. Sulphur
- b. Carbon
- c. Silica
- d. Tin

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Q17 100% Made In Star

Correct Answer:

100% 100% 100% 100%

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## Chemistry

Time Remaining 47/45 (Minutes)

Q.10

Test 5 Solids

CHEMISTRY NMDCAT

**The combined state of the metal is called?**

- a. Solid
- b. Metal
- c. Both A & B
- d. Minerals

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

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Time Remaining 42/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

**A solid, liquid, and gas can exist together at the**

- a. sublimation point
- b. triple point
- c. boiling point
- d. freezing point

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Time Remaining 47:45 (Minutes)

10:30

Test 5 Solids

CHEMISTRY NMDCAT

The glue comes under the example of

- a. crystalline solids
- b. amorphous solids
- c. simple solids
- d. compound solids

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Time Remaining 42/45 (Minutes)

Q11

Test 5 Solids

CHEMISTRY NMDCAT

The ions of the ionic crystals become free when it is in

- a. solid state
- b. compound state
- c. molten state
- d. none of above

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Correct Answer:

● 4 ● 11 ● 12 ● 13

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Time Remaining: 42/45 (Minutes)

Q.12

Test 5 Solids

CHEMISTRY NMDCAT

The crystal which is used to study Avogadro's number is called

- a. LiF
- b. NaI
- c. NaCl
- d. KCl

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Q.12 100% Made In Star Institute

Correct Answer:

4

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Time Remaining 42/45 (Minutes)

Q.11

Test 5 Solids

CHEMISTRY NMDCAT

The bond distance in the iodine molecule in solid iodine is

- a. 271.5pm
- b. 266.6pm
- c. 250pm
- d. 230pm

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Time Remaining 41/45 (Minutes)

Q38

Test 5 Solids

CHEMISTRY NMDCAT

The energy which is released when 1 mole of the ionic crystal is formed is known as

- a. lattice energy
- b. heat energy
- c. molar energy
- d. kinetic energy

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Correct Answer:

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Time Remaining 41/45 (Minutes)

Q.38

Test 5 Solids

CHEMISTRY NMDCAT

The isomorph of sodium fluoride is

- a. chlorine
- b. magnesium oxide
- c. sodium
- d. Sulphur

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Time Remaining 41/45 (Minutes)

Q.38

Test 5 Solids

CHEMISTRY NMDCAT

Lattice points have another name which is called lattice

- a. sites
- b. arrangements
- c. circles
- d. array

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Time Remaining 41/45 (Minutes)

Q37

Test 5 Solids

CHEMISTRY NMDCAT

The electrical properties of solid iodine include that it is

- a. non conductor
- b. conductor
- c. poor conductor
- d. moderate conductor

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Correct Answer:

1 2 3 4 5 6 7 8 9 10

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Time Remaining 41/45 (Minutes)

Q.38

Test 5 Solids

CHEMISTRY NMDCAT

**Electrical and thermal properties for some crystalline solids depend upon**

- a. surface
- b. area
- c. direction
- d. density

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1 2 3 4 5 6 7 8 9 10

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Time Remaining 41/45 (Minutes)

Q.38

Test 5 Solids

CHEMISTRY NMDCAT

**Covalent crystalline solids are soluble in**

- a. polar solvents
- b. non polar solvents
- c. water
- d. Normal saline

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Correct Answer:

☒ a ☐ b ☐ c ☐ d

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Time Remaining 41/45 (Minutes)

Q.40

Test 5 Solids

CHEMISTRY NMDCAT

The depressions among two layers of metals are also known as

- a. voids
- b. holes
- c. window
- d. shuttle

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Correct Answer:

☒ a ☐ b ☐ c ☐ d

Submit Quiz

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# TEST 5 SOLIDS

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## Q. 1

A crystal 'x' has very high melting point and is totally insoluble in water and does not conduct electricity is likely to be \_\_\_\_\_ crystal:

- a. Ionic
- b. Covalent
- c. Molecular
- d. Metallic

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Q. 2

Dry ice is solid:

a.  $SO_2$

b.  $CO$

c.  $CO$

d.  $O_2$

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## Q. 3

The inter ionic distance is maximum in the crystal lattice of

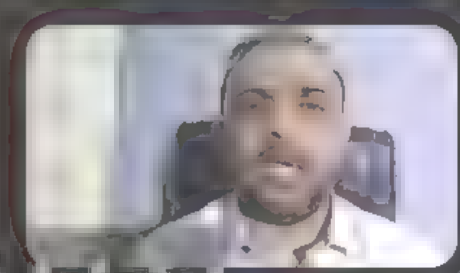
a. LiCl

b. NaCl

☒ c. RbCl

d. KCl

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## Q. 4

Among the properties which property correctly match with metallic solids:

- i. They show metallic lusters
- ii. Their conductance decreases with the rise of temperature
- iii. Malleable and ductile
- iv. Cubic and hexagonal close packing

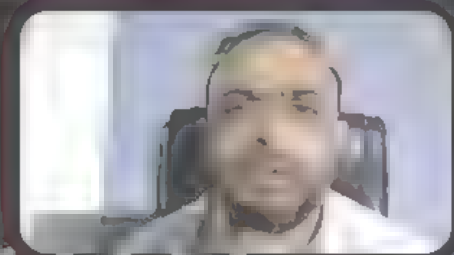
a. i, ii and iv

b. i, ii and iii

c. i, ii, iii and iv

d. i, iii and iv

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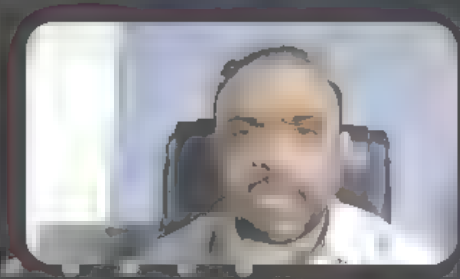


## Q. 5

Which statement about covalent solids is incorrect?

- a. They contain a network of atoms
- b. They have high melting points
- c. They are very hard and greater energy is required to break them
- d. **Their volatility is very high**

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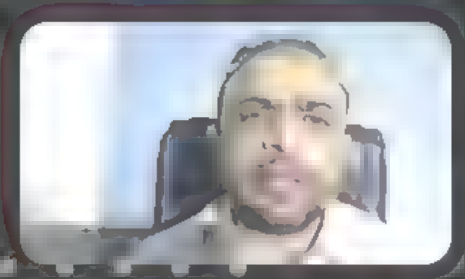


## Q. 6

Covalent crystals are bad conductor of electricity due to absence of free electrons and ions except:

- a. Silicon Carbide
- b. Graphite
- c. Cadmium iodide
- d. Born nitride

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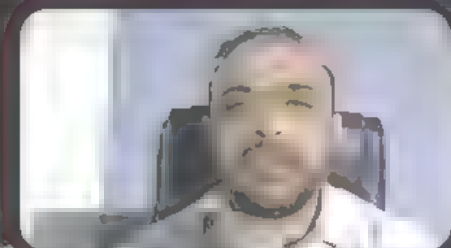
Q. 7

The covalent crystals having giant molecules like diamond and silicon carbide are:

- a. Soluble in all the solvents
- b. Insoluble in all the polar solvents only
- c. Soluble in all the non-polar solvents only

**d. Insoluble in all the solvents**

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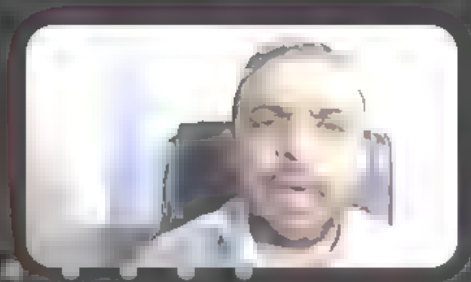


Q. 8

In diamond each carbon atom is:

- a.  $sp$ - hybridized
- b.  $sp^2$ -hybridized
- c.  $sp^3$ -hybridized
- d. unhybridized

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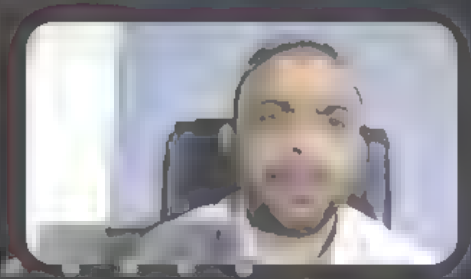


Q. 9

The overall structure of diamond looks like:

- ☐ a. Body centered cubic
- ☒ b. Face centered cubic
- c. End centered
- d. none of given

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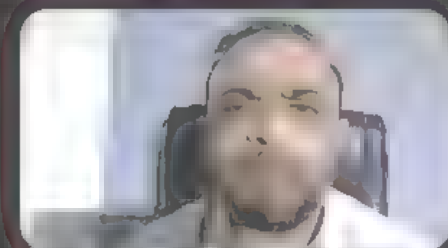


## Q. 10

Which one of the following pairs contains polar molecular solids?

- a. Iodine and Sugar
- ☒ b. Carbon dioxide and Ice
- c. Phosphorus and Carbon dioxide
- ☒ d. Sugar and Ice

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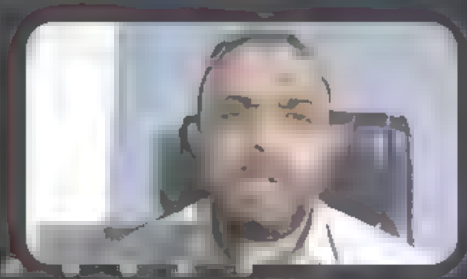
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## Q. 11

Valence bond theory treat metallic bond as:

- a. A coordinate covalent bond
- b. A localized covalent bond
- c. A de-localized covalent bond
- d. Van der Waal's force

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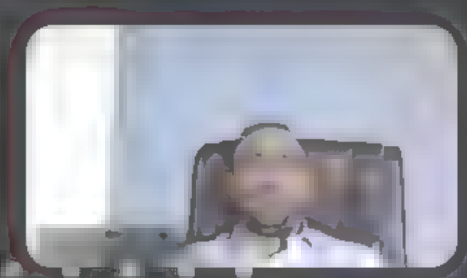


## Q. 12

Silicon carbide forms

- (a) ionic solid crystal (b) covalent crystal  
(c) molecular crystal (d) b, c

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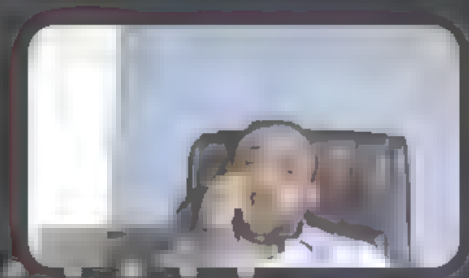
## Q. 13

Substance having m.p higher than  $500^{\circ}\text{C}$  and insoluble in  $\text{H}_2\text{O}$  and organic solvent and conductor in solid and phase:

a. Copper  
c. Silica

b. Sodium chloride  
d. Cell

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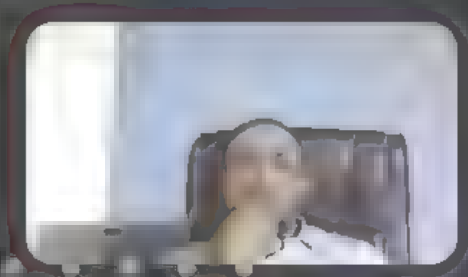


## Q. 14

The crystalline solid which show very slow chemical reactions

- (a) ionic
- (b) covalent
- (c) molecular
- (d) b, c

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**Q. 15**

The arrangement ABC, ABC .... is referred as

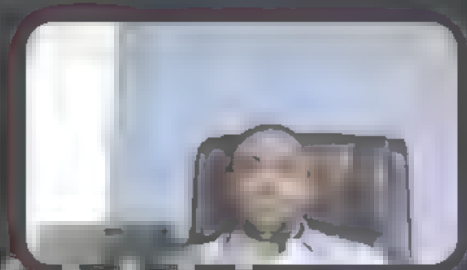
(a) cubic close packing

(b) octahedral closed packing

(c) hexagonal closed packing

(d) tetrahedral closed packing

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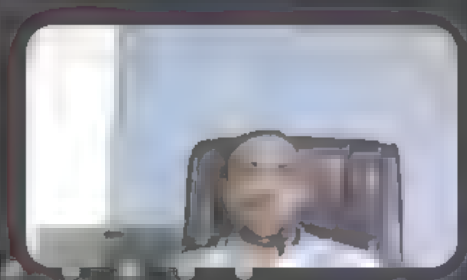


## Q. 16

Which statement is not true about metallic solid?

- (a) They are ductile and malleable
- (b) They are conductor of electricity
- (c) their compressibility increases
- (d) they are lustrous

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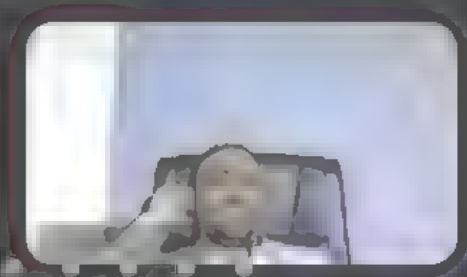


## Q. 17

Which solid does not contain covalent bonds

- (a) Copper
- (b) Diamond
- (c) Graphite
- (d) Ice

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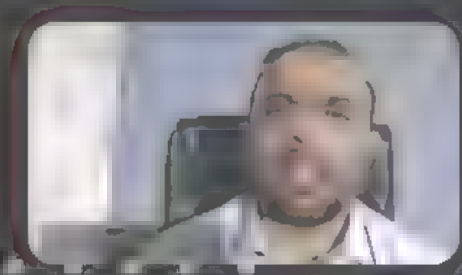


## Q. 18

The branch of science which deals with the study of the structure of crystal is called

- (a) Chymography
- (b) Crystallography
- (c) Chromatography
- (d) Industrial chemistry

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## Q. 19

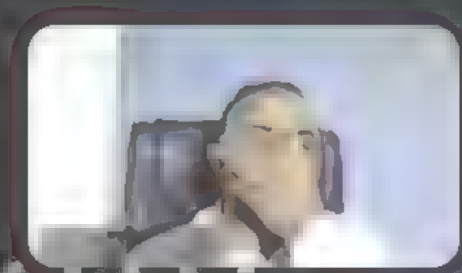
The diagram shows part of the lattice structures of solids X and Y.

What are the types of bonding present in X and Y?



	X	Y
A	covalent	metallic
B	ionic	covalent
C	ionic	metallic
D	metallic	ionic

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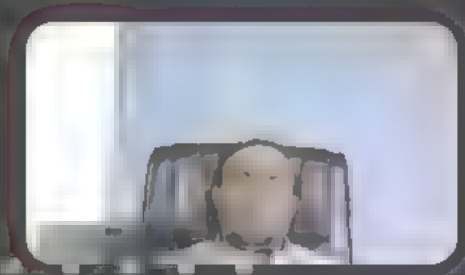


Q. 20

The crystalline part of other wise amorphous solids is called

- (a) crystal system      (b) crystallite  
(c) crystal lattice      (d) allotrope

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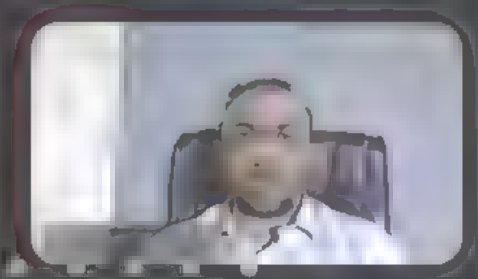


Q. 21

Which substance has diffused melting point

- (a) crystalline solid      (b) amorphous solid  
(c) metallic solid      (d) covalent solid

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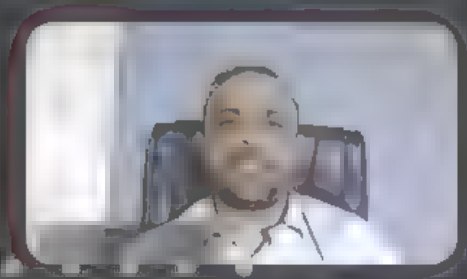


Q. 22

Which has the strongest bonding in the solid state?

- A. Hydrogen Chloride (HCl)
- B. Chlorine (Cl<sub>2</sub>)
- C. Xenon (Xe)
- D. Sodium Chloride (NaCl)

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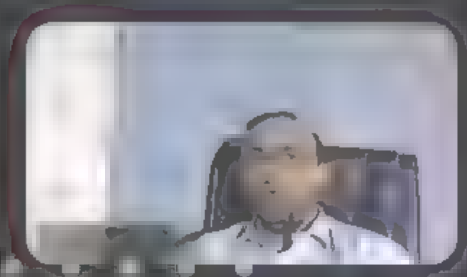


## Q. 23

When substance moves from a solid to a liquid state all of the following changes occur except

- A. Molecules become more disordered
- B. ~~Initial the molecule decreases~~
- C. Intermolecular forces become weaker
- D. Molecule become further separated

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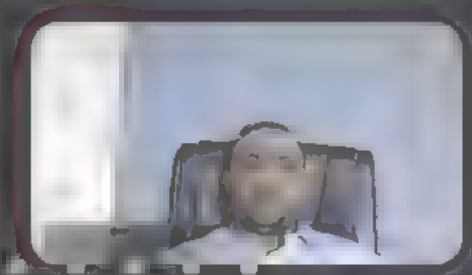


## Q. 24

When the atoms of the third layer are arranged in such a way that they directly lie above the atoms of the first layer then this arrangement is called

- A. ABAB (hexagonal)
- B. ABCABC (Cubic)
- C. Orthorhombic
- D. Rhombohedral

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Q.25

Which one is a conductor but is not malleable?

- A. Iron
- ☒ B. Graphite
- C. Silver
- D. Platinum

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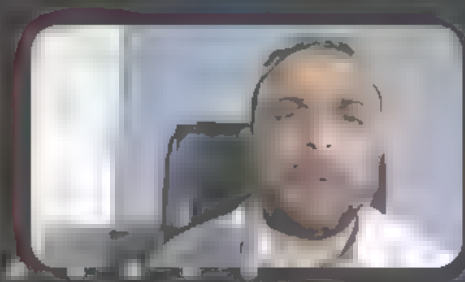


## Q. 26

A malleable solid is one which can be

- A. Converted into wire
- B. Converted into thin sheets
- C. Melted easily
- D. All of the above

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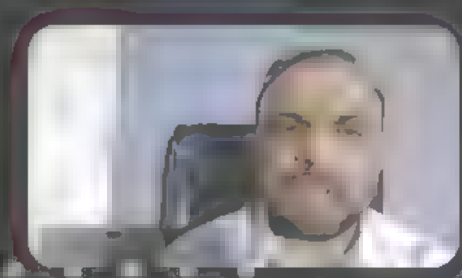


Q. 27

Buckyballs is an allotropic form of

- A. Sulphur
- B. Carbon
- C. Silica
- D. Tin

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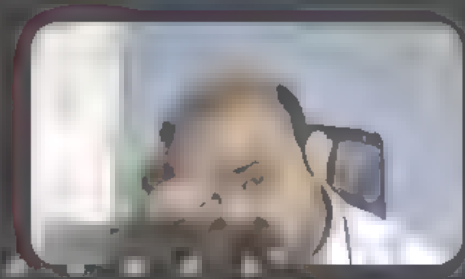
## Q. 28

The combined state of the metal is called?

- A. Solid
- B. Metal
- C. Both A & B

☒ Mineral

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## Q. 29

A solid, liquid, and gas can exist together at the

A. sublimation point

B. triple point

C. boiling point

D. freezing point

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## Q. 30

The glue comes under the example of

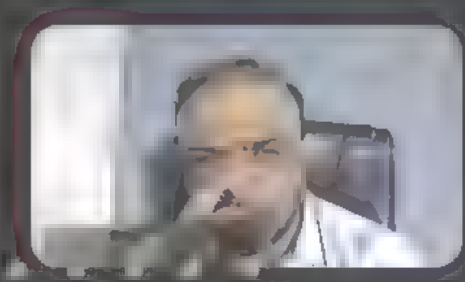
A. crystalline solids

☒ B. amorphous solids

C. simple solids

D. compound solids

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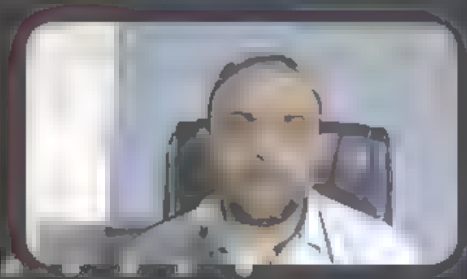


## Q. 31

The ions of the ionic crystals become free when it is in

- a) solid state
- b) compound state
- c) molten state
- d) none of above

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## Q. 32

The crystal which is used to study Avogadro's number is called

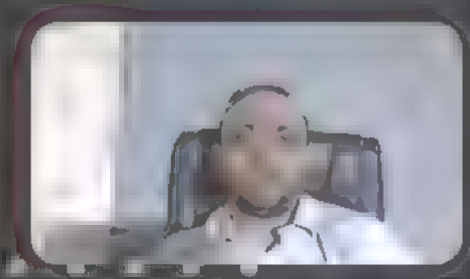
a)  $\text{NaCl}$

b)  $\text{NaI}$

c)  $\text{NaCl}$

d)  $\text{KCl}$

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## Q. 33

The bond distance in the iodine molecule in solid iodine is

a) 271 pm

b) 266.6 pm

c) 250 pm

d) 230 pm

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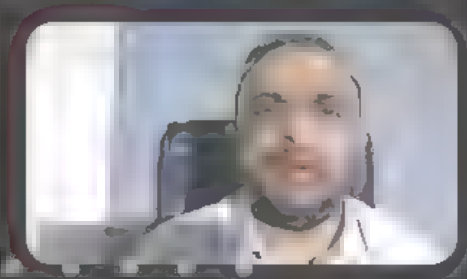


## Q. 34

The energy which is released when 1 mole of the ionic crystal is formed is known as

- a) lattice energy
- b) heat energy
- c) molar energy
- d) kinetic energy

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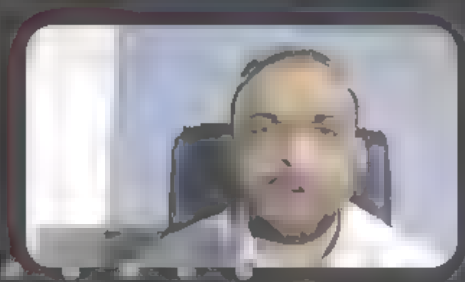


## Q. 35

The isomorph of sodium fluoride is

- a) chlorine
- b) magnesium oxide
- c) sodium
- d) Sulphur

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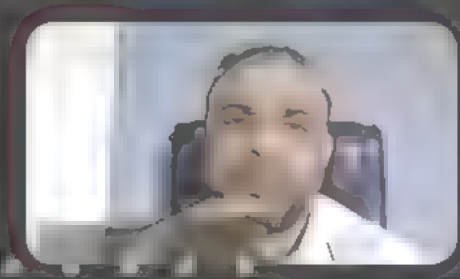


## Q. 36

Lattice points have another name which is called lattice

- a) ☒ lattice
- b) arrangements
- c) circles
- d) array

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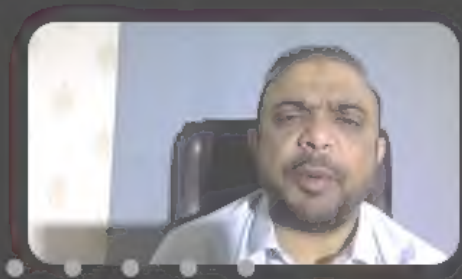


## Q. 37

The electrical properties of solid iodine include that it is

- a) non conductor
- b) conductor
- c) poor conductor
- d) moderate conductor

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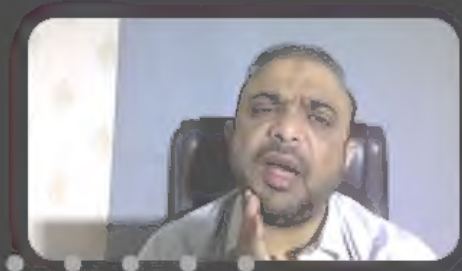
سید اختر عباس جعفری screen

## Q. 38

Electrical and thermal properties for some crystalline solids depend upon

- a) surface
- b) area
- c) direction
- d) density

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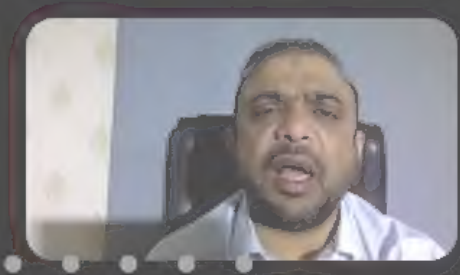


## Q. 39

Covalent crystalline solids are soluble in

- a) polar solvents
- b) non polar solvents
- c) water
- d) Normal saline

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## Q. 40

The depressions among two layers of metals are also known as

- a) voids
- b) holes
- c) window
- d) shuttle

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